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An Explanation~~

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lab v2Predicting

Products of Double

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How to Predict Products

of Chemical Reactions |

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How to Pass Chemistry

Solubility Rules and

How to Use a Solubility

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part 2 Titration

Experiment \u0026

Calculate the Molarity

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Compounds | How to

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Stoichiometry Using

Molarity as a

Conversion Factor |

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Oxidation and

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Reduction (Redox)

Reactions Step-by-Step

Example Writing and

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Predicting Products

Transcription and

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DNA Replication:

Copying the Molecule

of Life What is DNA?

How to Find Limiting

Reactants | How to

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Net Ionic Equation

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Introduction to Double  
Replacement Reactions  
Predicting Products of  
Single Replacement  
Reactions Types of  
Chemical Reactions

After watching this,  
your brain will not be  
the same | Lara Boyd |  
TEDxVancouver

---

Classifying Chemical  
Reactions - Double  
Replacement

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Linguistics, Style and



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Writing in the 21st  
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Replacement Reactions.  
Objectives. The

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objectives of this lab are to: a) Perform and observe the results of a variety of double replacement reactions, b) Become familiar with some of the observable signs of these reactions, c) Identify the products formed in each of these reactions, d) Write balanced chemical equations for each double replacement

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The objectives of this lab are to: a) Perform and observe the results of a variety of double replacement reactions, b) Become familiar with some of the observable signs of these reactions, c) Identify the products formed in each of these

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reactions, d) Write balanced chemical equations for each double replacement reaction studied.

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Double Replacement  
Reactions A double  
replacement reaction  
has the form:  $AB + CD$

$AD + CB$  There are

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## Double

four different possible  
outcomes to a reaction  
such as this: [1]

Formation of a gas.

There are certain  
compounds which are  
unstable and decompose  
to water and a gas.

Three common ones are  
 $\text{H}_2\text{CO}_3$ ,  $\text{H}_2\text{SO}_3$  and  
 $\text{NH}_4\text{OH}$ . They  
decompose like this:  $\text{H}_2\text{CO}_3 \rightarrow \text{H}_2\text{O} + \text{CO}_2$

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Double Replacement  
Reactions All double  
replacement reactions  
have the general form:  
 $AB + CD \rightarrow AD + CB$   
Reactions that can be  
classified as double  
replacements include

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Double

precipitation reactions,  
neutralization reactions  
and gas forming  
reactions.

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10: Double

Replacement Reactions

(Experiment) -

Chemistry ...

The general form of a

double-replacement

(also called double-

displacement) reaction

is: (11.9.1)  $AB + CD$



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## Double

**AD + BC** In this reaction, A and C are positively-charged cations, while B and D are negatively-charged anions. Double-replacement reactions generally occur between substances in aqueous solution.

11.9: Double  
Replacement Reactions  
- Chemistry LibreTexts

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Acid-Base – Also called  
neutralization reactions.

These double  
replacement reactions  
occur when an acid and  
a base react to make a

salt and water.  $HA +$   
 $BOH \rightarrow BA + H_2O$ .

$HNO_3(aq) + NaOH$   
 $(aq) \rightarrow NaNO_3(aq) + H$   
 $2O(l)$  For today 's

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experiment you will  
classify reactions as one  
of the main 5 types of  
chemical reactions.

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Lab 6 Introduction |  
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Use your observations  
above, information in  
the introduction, and  
reading to classify the

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## Double

above chemical reactions in the table below. First: classify the reaction as a main type: either synthesis, decomposition, combustion, single replacement OR double replacement. If no reaction occurred, say N/A. Record your choice in the space provided.

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Double

Lab 6 Worksheet |

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Double Replacement  
Reactions Laboratory

Manual Answers In this lab, double replacement reactions between compounds were done in order to determine the equation and description of a new substance. During the lab, each participant

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was given drop bottles,  
spot plates.

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Golon \* Single

Replacement Reaction

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Lab Manual \* single

replacement reactions in  
part a of this lab we will  
examine single

replacement reactions

this is one type of  
oxidation reduction  
reaction or redox

reaction because it  
occurs via a transfer of  
electrons single

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