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Stoichiometry: Limiting Reactant, Left Over Excess Reactant, Percent Yield | Study

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Chemistry Stoichiometry Supplemental Study Guide Stoichiometry The atomic ratios in each compound are also the relative number of atomic mass units of its elements. The first example is nitrous oxide (N_2O), as shown in Table 1. The relative masses were obtained by multiplying the atomic ratios and atomic masses. Stoichiometry

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