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The measurement has two significant figures. The measurement has four significant figures. Decimal point is present, so these zeros are significant. The rules are summarized in the following flowchart:

qq Nonzero digits 3 4 0
0 g of sulfur pp Nonzero digits 4 0
0 0. mL of oxygen pp

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CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 8.

Stoichiometry. So far in your chemistry course, you have learned that chemists count

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Answer Key quantities of elements and compounds in terms of moles and that they relate moles of a substance to mass by using the molar mass. In addition, you have learned to write chemical equations so that they represent the rearrangements of atoms that take place during chemical reactions, and you have learned to balance these equations.

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What is the molarity of a solution that is made by diluting 50.00 mL of a 4.74 M ...

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PROBLEM SOLVING 1. Determine
the percentage composition of
each of the following compounds:
a. sodium oxalate, $\text{Na}_2\text{C}_2\text{O}_4$
ans: 34.31% Na, 17.93% C,
47.76% O b. ethanol, $\text{C}_2\text{H}_5\text{OH}$
ans: 52.13% C, 13.15% H, 34.72%
O c. aluminum oxide, Al_2O_3
ans: 52.92% Al, 47.08% O d.
potassium sulfate, K_2SO_4 ans:
44.87% K, 18.40% S, 36.72% O 2.

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CHEMFILE MINI-GUIDE TO
PROBLEM SOLVING General Plan
for Solving Molarity Problems

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Answer Key
Mass of solute in g
1 Amount of solute in mol
M moles solute liter solution
2 Volume of solution in L
3 Molar concentration, M
4 Convert using the molar mass of the solute.
MOLARITY Molarity is the most common way to express concentration in chemistry.

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CHEMFILE MINI-GUIDE TO
PROBLEM SOLVING Neither pressure nor temperature is constant; therefore, the combined gas law must be used to calculate the volume of CO₂ at the final temperature and pressure.
multiply by the volume ratio, CO₂
2 use the combined gas law to convert from the initial

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temperature and pressure to the
final temperature and pressure C
2 H 4

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Neither
pressure nor temperature is
constant; therefore, the combined
gas law must be used to calculate
the volume of CO₂ at the final
tem-perature and pressure.

multiply by the volume ratio, CO
2 use the combined gas law to
convert from the

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PROBLEM SOLVING 1. Carry out the following computations, and express the result in scientific notation: a. $2.290 \times 10^7 \text{ cm} \times 4.33 \times 10^3 \text{ s}$ ans: $5.29 \times 10^3 \text{ cm/s}$ b. $1.788 \times 10^5 \text{ L} \times 7.111 \times 10^3 \text{ m}^2$ ans: $2.514 \times 10^3 \text{ L/m}^2$ c. $5.515 \times 10^4 \text{ L} \times 6.04 \times 10^3 \text{ km}$ ans: 9.13 L/km d. $3.29 \times 10^4 \text{ km} \times 1.48 \times 10^2 \text{ min}$ ans: $2.22 \times 10^2 \text{ km/min}$

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