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How to Pass

Chemistry **SN1, SN2,**

E1, \u0026 E2

Reaction

Mechanism Made

Easy! Oxidation and

Reduction (Redox)

Reactions Step-by-

Step Example

Solubility Rules and

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Solubility Table

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~~TUTOR HOTLINE~~

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~~Moles Using Molar~~

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~~Metric System \u0026~~

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shrags. Terms in this

set (114) Avogadro's

Constant. $6.022 \times$

10^{23} mol^{-1} . Speed

of Light formula . $c =$

$3.00 \times 10^8 \text{ m/s}$ $c = ?$

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Planck's Constant.

$$h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$$

$$E_{\text{photon}} = hf$$

Rydberg Constant.

$$R_H = 2.179 \times 10^{-18}$$

$$\text{J} \quad E = -R_H \left(\frac{1}{n_f^2} - \frac{1}{n_i^2} \right)$$

Choose the
three ...

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science that deals
with the materials of
the universe and the
changes these
materials undergo;
the understanding of
solids, liquids, and
gases; a central

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observations;
formulate hypotheses;
perform experiments .
the three main ...

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Key Concepts: Terms
in this set (61)
Molarity. Moles of
Solute/Liters of
Solution. Moles of
Solute. Molarity x
Volume of Solution.
Volume of Solution.
Moles of
Solute/Molarity. What
is the oxidation
number of chlorine in

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ClO3-? a. -1 b. 0 c. -6
d. +5 e. none of ...

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The final exam is cumulative and will assume you have reviewed chapters 1-7 with greater emphasis placed on material after chapter 5. (for chapters 1-5, look at previous midterm exam review guides)

Chapter 6:

Thermochemistry a)

Be able to state and

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this set (58) Ionic

Compound. a compound composed of positive and negative ions that are combined so that the numbers of positive and negative charges are equal. Covalent

Compound. a chemical compound formed by the sharing of electrons . The Ideal Gas Law. $PV =$

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101 c. If you want
your scores posted,
mark A under the
option column 2. Use
a #1 or #2 pencil for
marking the answer
sheets. Fill in the
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completely. 3. DO
NOT ...

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calculations.

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answers CHEM 101
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together) The final
exam is cumulative
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